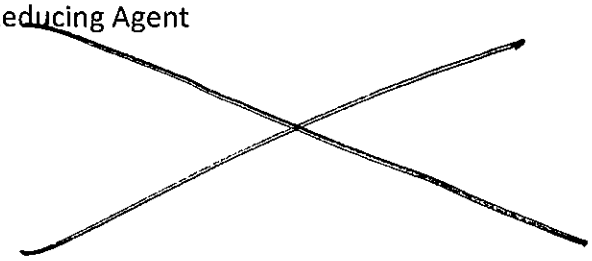
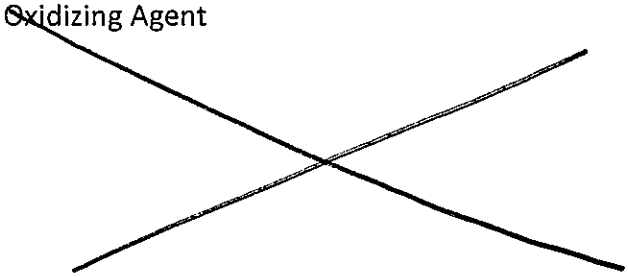
Unit #10: Vocabulary (definitions are on Miss Virga's website)

Anode electrode where oxidation occurs "AN OX"	Cathode electrode where reduction occurs "RED CAT"
Electrochemical Cell two types - voltaic & electrolytic	Electrode conductive surfaces (metals) where oxidation or reduction occur
Electrolytic Cell cell that uses electrical energy to force a nonspontaneous chemical rxn to occur	Redox reactions that involve the transfer of ELECTRONS
Half Reactions reduction: $Cu^{2+} + 2e^- \longrightarrow Cu$ oxidation: $Fe \longrightarrow Fe^{3+} + 3e^-$	Reducing Agent 
Oxidizing Agent 	Oxidation Number "charge" that can be assigned to atoms in elements, compounds, or ions
~~Galvanic~~ or Voltaic Cell cell that spontaneously converts chemical energy into electrical	Salt Bridge necessary to allow ions to flow in electrochemical cells to prevent charge build up