

Unit #7: Solution Vocabulary

<p>Colligative Properties</p> <p>boiling point <u>ELEVATION</u> (↑)</p> <p>freezing point <u>DEPRESSION</u> (↓)</p>	<p>Concentration</p> <p>a measure of how much solute is dissolved in a specific amount of solvent or solution</p>
<p>Electrolyte vs Nonelectrolyte</p> <p>↓ substance that can conduct electricity when dissolved in H₂O</p>	<p>Molarity</p> <p>moles of solute per liter of solution</p>
<p>Parts Per Million (ppm)</p> <p>Table T</p> $\frac{\text{mass solute}}{\text{mass solution}} \times 1,000,000$	<p>Saturated Solution</p> <p>contains the maximum amount of dissolved solute for a given amount of solvent at a specific temp/pressure</p>
<p>Supersaturated Solution</p> <p>contains <u>more</u> dissolved solute than a saturated solution at the same temp.</p>	<p>Unsaturated Solution</p> <p>contains less dissolved solute for a given temperature and pressure than a saturated solution</p>
<p>Solubility</p> <p>↓ the maximum amount of solute that will dissolve in a given amt of solvent</p>	<p>Solute</p> <p>↓ one or more substances dissolved in a solution</p>
<p>Soluble vs Insoluble</p> <p>↓ can dissolve ↓ cannot dissolve</p>	<p>Solvent</p> <p>the substance that dissolves a solute to form a solution</p>
<p>"Like Dissolves Like"</p> <p>polar solutes dissolve in polar solvents</p> <p>(nonpolar solutes dissolve in NP solvents)</p>	<p>Precipitate</p> <p>a <u>solid</u> produced during a chemical rxn in a solution</p>
<p>Dilute vs Concentrated</p> <p>↓ a little solute dissolved</p> <p>↓ A LOT of solute dissolved</p>	<p>Aqueous</p> <p>a solution in which the solvent is water</p>